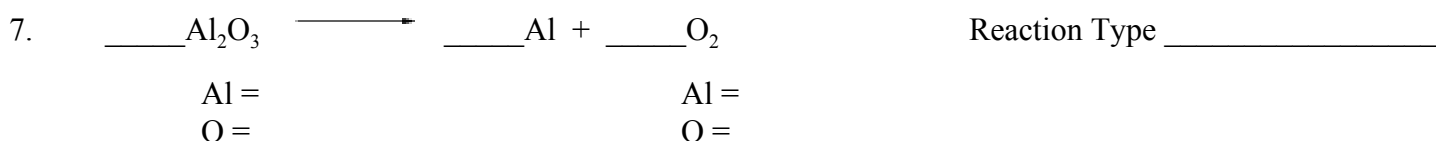
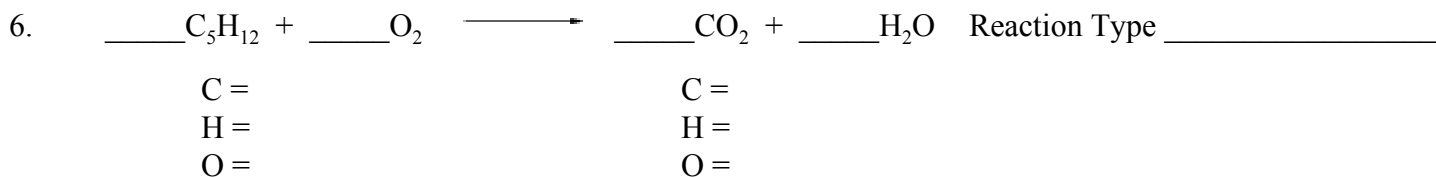
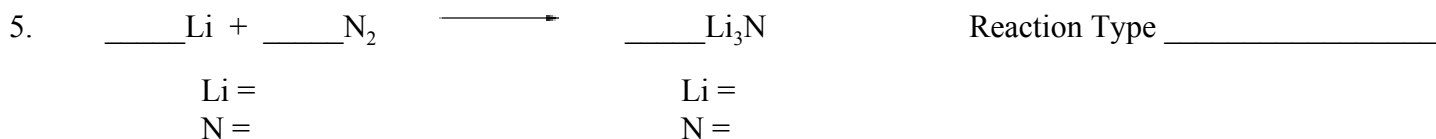
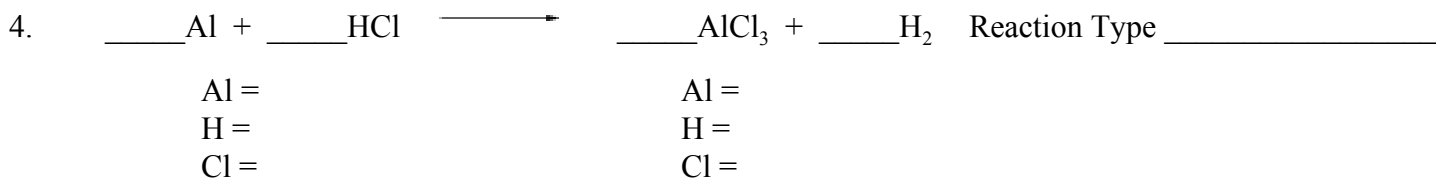
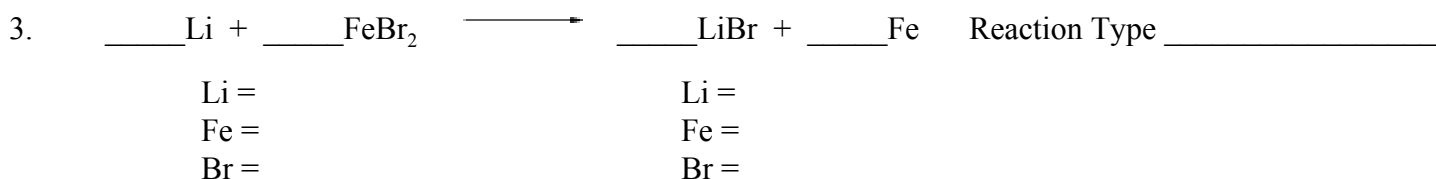
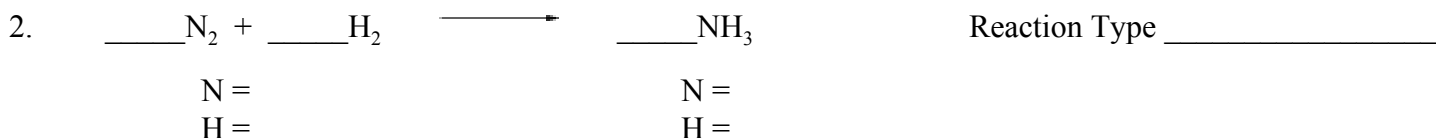
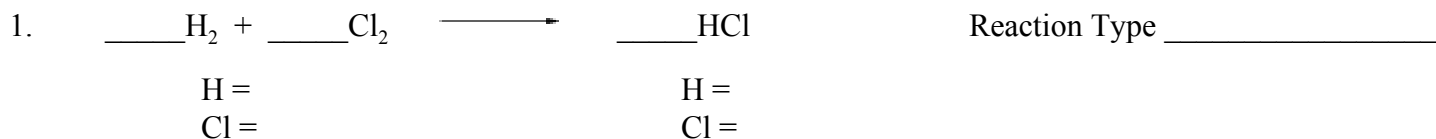


The Periodic Table and Bonding WS 4

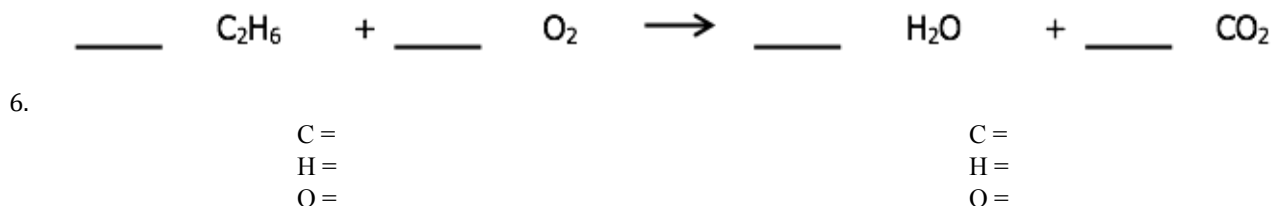
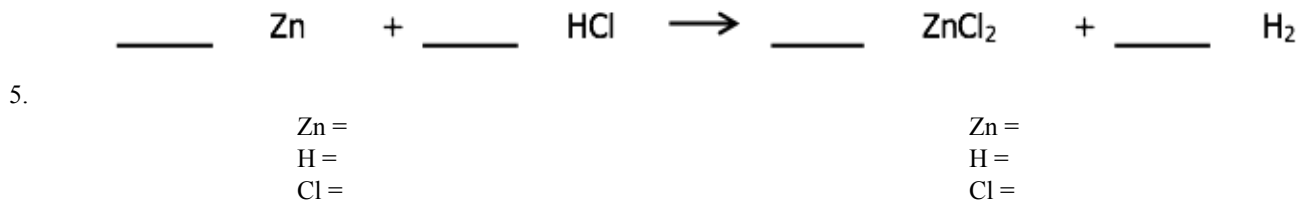
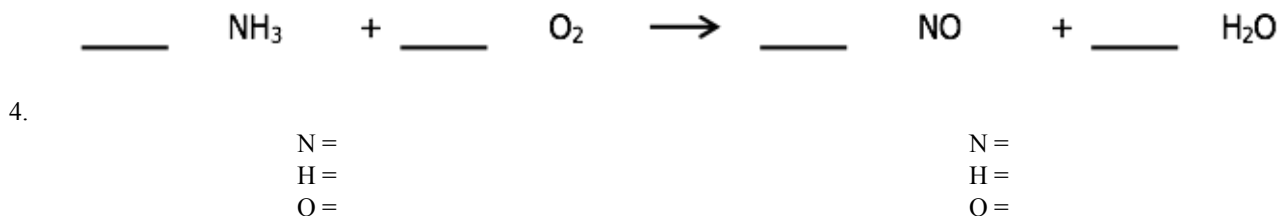
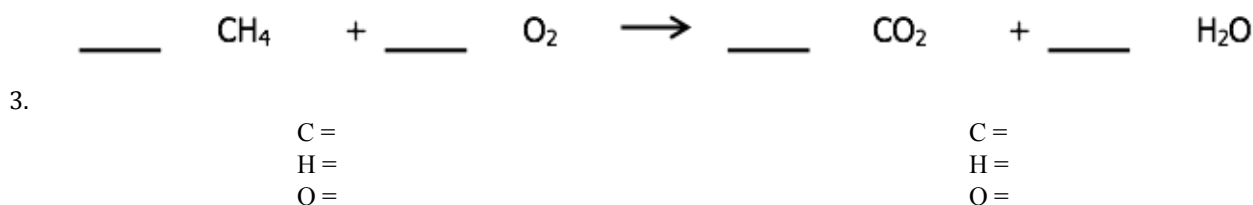
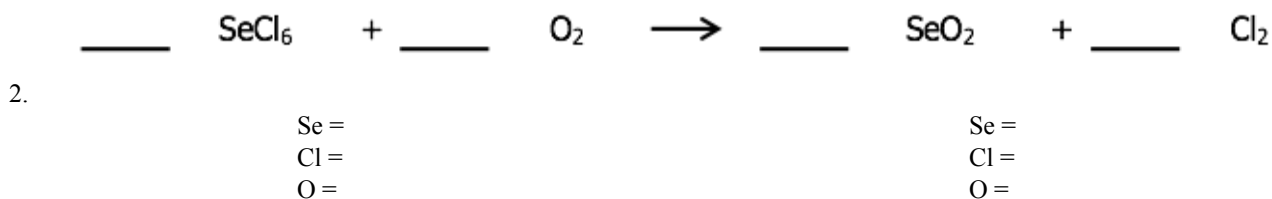
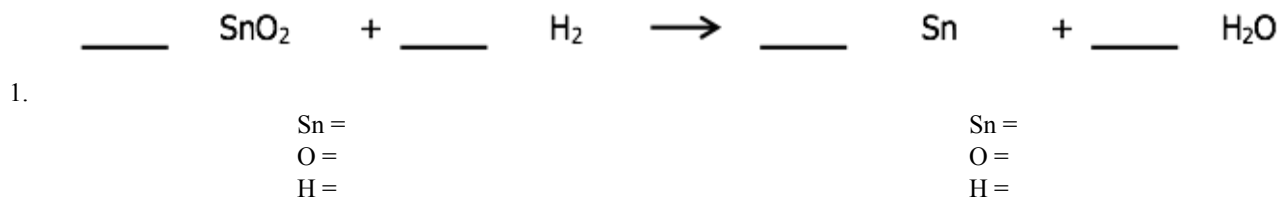
Law of Constant Mass states that matter cannot be created or destroyed, only changed. This means that in a chemical reaction the amount of reactant you start with must equal the amount of product that you end with. This requires a chemical equation to be balanced.

Balancing chemical equations involves counting the atoms of each element and making sure that they are equal on both sides. If they are, the equation is said to be “balanced.” If not, then the equation must be balanced by adding coefficients. **NEVER** change the subscript number of a formula when balancing, this changes the compound.



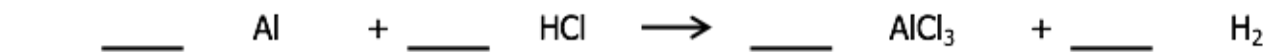
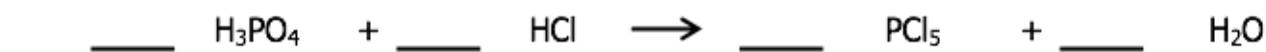
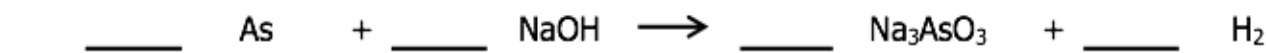
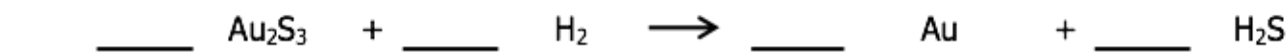
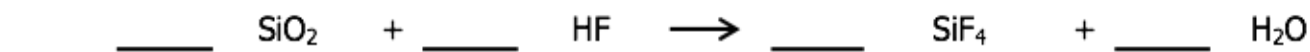
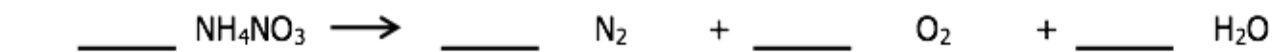
The Periodic Table and Bonding WS 5

For each of the following chemical reactions place the number the correctly balances the equation.



PHYSICAL SCIENCE

Name: _____

DSHS
Mrs. EllisAl =
O =Al =
O =Al =
H =
Cl =Al =
H =
Cl =H =
P =
O =
Cl =H =
P =
O =
Cl =As =
Na =
O =
H =As =
Na =
O =
H =Au =
S =
H =Au =
S =
H =Si =
O =
H =
F =Si =
O =
H =
F =N =
H =
O =N =
H =
O =